General Chemistry II (CHM 2046)

Potentially useful information, Exam 1

\[ K = \degree C + 273.15 \]
\[ N_A = 6.022 \times 10^{23} \]
\[ 1 \text{ J} = 1 \text{ kg} \cdot \text{m}^2/\text{s}^2 \]
\[ 1 \text{ cal} = 4.184 \text{ J} \]
\[ R = 0.08206 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}} \]
\[ 1 \text{ atm} = 101.325 \text{ kPa} = 760 \text{ mm Hg} = 760 \text{ torr} = 29.921 \text{ in Hg} = 1.01325 \text{ bar} \]
\[ q = s \cdot m \cdot \Delta T \]
\[ q = \Delta H_m \]
\[ E = \frac{hc}{\lambda} = h \nu \]
\[ h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s} \]
\[ c = 3.00 \times 10^8 \text{ m/s} \]
\[ 1 \text{ eV} = 1.602 \times 10^{-19} \text{ J} \]

Write your name, then bubble your n-number (xxxxxxxxx) leaving a blank at the end, and test form letter (after starting exam, see next page) on the RED ParSCORE form.
Bubble in your answers on the RED ParSCORE. Bubble both A and B to indicate AB. No extra time will be given to transfer answers to the Scantron sheet. You may remove the top sheet of the exam after you have been given instructions to begin the exam.

Do not turn or remove this page until you are told to begin the exam.